

# Number of Atoms by Formula

Use the formulas to determine how many atoms are in each molecule. Use the terms in the word box to label each molecule.

ferric oxide  
mercurous chloride  
potassium carbonate  
ethane

hydrogen peroxide  
sulfuric acid  
calcium chloride  
sodium sulfide

sodium chloride  
phosphoric acid  
ammonium bromide  
phosphocreatine

Example:  $\text{CO}_2$   
Atoms: 1 atom of carbon and two atoms of oxygen = 3 atoms  
Name: carbon dioxide

<p><b>1</b> <math>\text{NaCl}</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>2</b> <math>\text{H}_2\text{O}_2</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>3</b> <math>\text{Hg}_2\text{Cl}_2</math></p> <p>Atoms: _____</p> <p>Name: _____</p>
<p><b>4</b> <math>\text{Fe}_2\text{O}_3</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>5</b> <math>\text{H}_3\text{PO}_4</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>6</b> <math>\text{K}_2\text{CO}_3</math></p> <p>Atoms: _____</p> <p>Name: _____</p>
<p><b>7</b> <math>\text{CaCl}_2</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>8</b> <math>\text{NH}_4\text{Br}</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>9</b> <math>\text{C}_2\text{H}_6</math></p> <p>Atoms: _____</p> <p>Name: _____</p>
<p><b>10</b> <math>\text{H}_2\text{SO}_4</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>11</b> <math>\text{Na}_2\text{SO}_3</math></p> <p>Atoms: _____</p> <p>Name: _____</p>	<p><b>12</b> <math>\text{C}_4\text{H}_{10}\text{N}_3\text{O}_5\text{P}</math></p> <p>Atoms: _____</p> <p>Name: _____</p>